



Oxford Cambridge and RSA

Tuesday 12 October 2021 – Morning

A Level Chemistry B (Salters)

H433/02 Scientific literacy in chemistry

Advanced Notice Article

Time allowed: 2 hours 15 minutes



INSTRUCTIONS

- Do **not** send this Advanced Notice Article for marking. Keep it in the centre or recycle it.

INFORMATION

- This is a clean copy of the Advance Notice Article you have already seen.
- This document has **8** pages.

Smartphones: Smart Chemistry

Adapted from an article by Brian Rohrig in the April/May 2015 ChemMatters online magazine of the American Chemical Society

Could you last a day without your cell phone? As many as 84% of U.S. residents could not, according to a recent poll conducted by *Time* magazine. It is hard to believe that 20 years ago, hardly anyone even owned a cell phone. And now the cell phone has morphed into something bigger and better – the smartphone. Worldwide, more than *one billion* smartphones were purchased last year. If you own a smartphone, you are probably aware that in a year or two, it will be practically obsolete, because the smartphone just keeps getting smarter.

In the 1950s, you would have needed a whole bank of computers on an entire floor of an office building to do what you are able to do with a single smartphone today. Even a low-end smartphone has more computing power than the computer system the National Aeronautics and Space Administration (NASA) used to put a man on the moon. Amazingly, you can surf the Internet, listen to music, and text your friends with something that fits in the palm of your hand. None of this would be possible without chemistry, and every time you use your smartphone, you are putting chemistry into action.

Smartphone chemistry

If you are wondering what chemistry has to do with smartphones, just look at the periodic table. Of the 83 stable (nonradioactive) elements, at least 70 of them can be found in smartphones! That's 84% of all stable elements.

Metals are what make smartphones so “smart.” An average smartphone may contain up to 62 different metals. One rather obscure group of metals – the rare-earth metals – plays a vital role. These rare-earth metals include scandium and yttrium, as well as elements 57–71, known as the lanthanides, because they begin with the element lanthanum. The lanthanides often appear as the first of two free-floating rows located at the bottom of the periodic table. Scandium and yttrium are included because their chemical properties are similar to those of the lanthanides.

If you examine several varieties of smartphones, you can find 16 of the 17 rare earth metals. The only one you will not find is promethium, which is radioactive.

Many of the vivid red, blue, and green colours you see on your screen are due to rare-earth metals, which are also used in the phone circuitry and in the speakers. Also, your phone would not be able to vibrate without neodymium ($[\text{Kr}]4d^{10}4f^45s^25p^65d^06s^2$) and dysprosium ($[\text{Kr}]4d^{10}4f^{10}5s^25p^65d^06s^2$).

Rare-earth metals are also used in many other high-tech devices. They are found in televisions, computers, lasers, missiles, camera lenses, fluorescent light bulbs, and catalytic converters. Rare-earth elements are so important in the electronics, communications, and defence industries that the U.S. Department of Energy dubbed them the “technology metals.”

Smartphone's display

The screen allows you to see the phone's display. If you have ever dropped your phone without damaging the screen, you were probably relieved. Smartphone screens are designed to be extremely tough.

This toughness is actually the result of a serendipitous accident. In 1952, a chemist at Corning Glass Works was trying to heat a sample of glass to 600 °C in a furnace when a faulty thermostat caused it to be heated to 900 °C. Upon opening the door, he was glad – and surprised – to find that his glass sample was not a melted pile of goo and that it had not ruined the furnace. When he took it out with tongs, he dropped it on the floor (another accident). But instead of breaking, it bounced!

Thus, was born the world's first synthetic glass-ceramic, a material that shares many properties with both glass and ceramic. Glass is an amorphous solid, because it lacks a crystalline structure (**Fig. 1(a)**). The structure is not ordered, but is arranged more like a liquid, yet frozen in place. Because glass does not contain planes of atoms that can slip past each other, there is no way to relieve stress. Excessive stress forms a crack, and bonds on the surface of the crack break. As the crack grows, the intensity of the stress increases, more bonds break, and the crack widens until the glass breaks.

Ceramics, on the other hand, tend to be regular and crystalline (**Fig. 1(b)**). When they form crystals, the strong force of attraction between the planes makes it difficult for one plane to slip past another. Ceramics are therefore brittle. They resist compression, but they can break when they are bent. The attraction between planes can be increased by the introduction of ions into the structure.

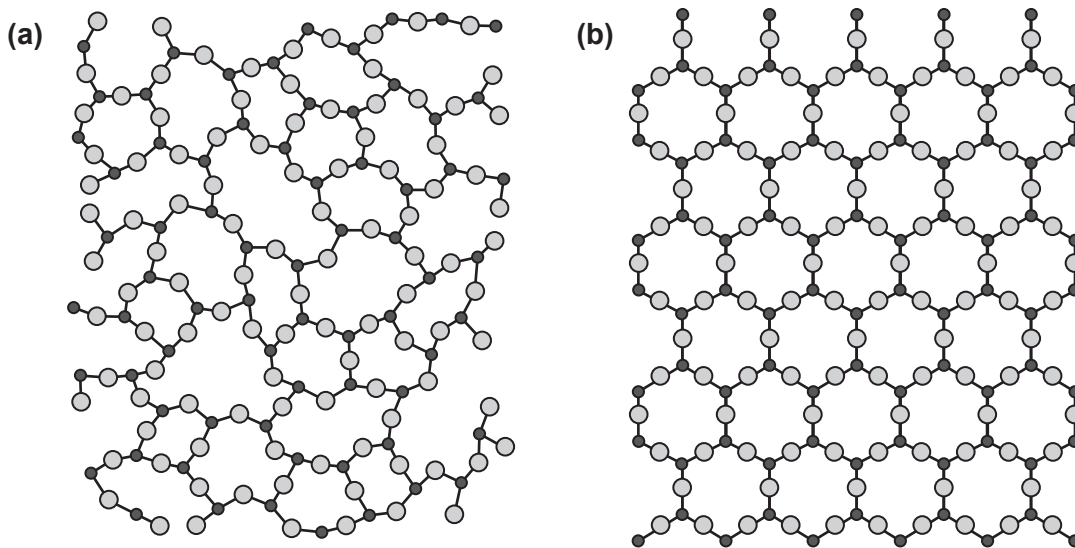


Fig. 1

Two dimensional comparison of the chemical structures of **(a)** an amorphous solid made of silicon dioxide (glass) and **(b)** a crystal of silicon dioxide (ceramic)

In three dimensions, the structures can be represented as in **Fig. 2**:

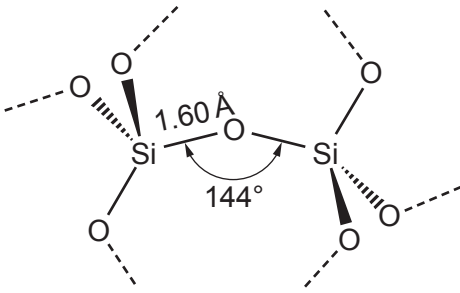


Fig. 2

The combination of glass and ceramic forms a material that is tougher and stronger than each of the materials by themselves. A glass-ceramic is formed by overheating the glass, so a portion of its structure is transformed into a fine-grained crystalline material. Glass-ceramics are at least 50% crystalline, and in some cases, they are more than 95% crystalline.

This amazing glass-ceramic material is so resistant to heat that it has been used in the nose cones of supersonic-guided missiles used by the military. As a result of the success of glass-ceramic materials, the Corning Glass Works Company undertook a large research effort to find ways to make ordinary transparent glass as strong as glass-ceramic products. By 1962, Corning had developed a very strong type of chemically strengthened glass, unlike anything ever seen before. This super-strong glass would eventually make its way to nearly every smartphone screen. It is so strong it goes by the name, **Gorilla Glass. Laboratory tests have shown that Gorilla Glass can withstand 100,000 pounds of pressure per square inch!**

Gorilla Glass is composed of oxygen, silicon and aluminium—also called aluminosilicate glass—along with sodium ions (**Fig. 3**).

But Gorilla Glass gains its tremendous strength through one final step, in which the glass is chemically strengthened. The glass is put into a molten bath of potassium salt, usually potassium nitrate, at 300 °C. Potassium ions replace sodium ions in the structure and take up more space in the glass than do sodium ions. The anions are aluminosilicate.

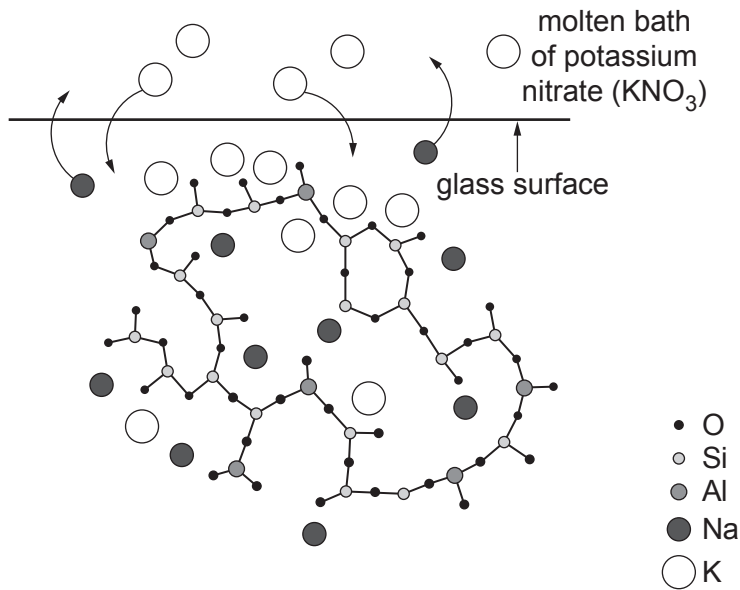


Fig. 3

Cramming larger ions into the spaces formerly occupied by smaller ions results in a compressing of the glass. Compressed glass is very strong. As a result of this compression, a lot of elastic potential energy is stored in the glass, much like the elastic potential energy that you might find in a compressed spring.

What's behind a touchscreen?

As every smartphone user knows, the screen on a smartphone is far more than just a tough piece of glass. It is a screen that responds to your touch—aptly named a touchscreen—giving you a personal connection to your phone.

There are two basic categories of touchscreens. The first category of touchscreens, called **resistive touchscreens**, can be touched with any type of material and they will still work. You can activate the screen even if wearing gloves. Resistive touchscreens are used when you sign for a delivery on a display screen.

Resistive touchscreens are composed of two thin layers of conductive material under the surface. When you press down a resistive touchscreen, it physically indents, causing the two layers to touch, completing the circuit and changing the electrical current at the point of contact. The software recognizes a change in the current at these coordinates and carries out the action that corresponds with that spot. Resistive touchscreens are also known as pressure-sensitive screens.

Smartphones use the second basic category of touchscreens, called **capacitive touchscreens** which are electrical in nature. A capacitor is any device that stores electricity.

When a finger presses down on a capacitive touchscreen, a very small electrical charge is transferred to the finger, creating a voltage drop on that point of the screen. A controller within the smartphone processes the location of this voltage drop and orders the appropriate action.

Glass, being an insulator, does not conduct electricity. Even though the glass contains ions, they are locked into place, stopping electricity from flowing through. So, the glass screen must be coated with a thin transparent layer of a conductive substance, usually indium tin oxide, which is laid out in crisscrossing thin strips to form a grid pattern. One form of indium tin oxide has percentages by mass: 53% In, 28% Sn and 19% O.

This conductive grid acts as a capacitor, storing very small electrical charges. When you touch the screen, a tiny bit of this stored electrical charge enters your finger—not enough for you to feel but enough for the screen to detect. As this electrical charge leaves the screen and enters your finger, the screen registers a voltage drop, the location of which is processed by the software, which orders the resulting action.

This tiny bit of electrical current enters your finger because your skin is an electrical conductor—primarily due to the combination of salt and moisture on your fingertips, creating an ionic solution. Your body actually becomes part of the circuit, as a tiny bit of electricity flows through you every time you use the touchscreen on your phone.

Smartphone technology is evolving at a dizzying pace. You can now use your smartphone to check your blood sugar, adjust your home's thermostat, and start your car. Twenty years ago, no one envisioned that people would someday take more pictures with their cell phones than with their stand-alone cameras. It is anyone's guess what will come next. Thanks to the intersection of chemistry and innovation, the possibilities are limitless.

END OF ADVANCE NOTICE ARTICLE

OCR
Oxford Cambridge and RSA

Copyright Information

OCR is committed to seeking permission to reproduce all third-party content that it uses in its assessment materials. OCR has attempted to identify and contact all copyright holders whose work is used in this paper. To avoid the issue of disclosure of answer-related information to candidates, all copyright acknowledgements are reproduced in the OCR Copyright Acknowledgements Booklet. This is produced for each series of examinations and is freely available to download from our public website (www.ocr.org.uk) after the live examination series.

If OCR has unwittingly failed to correctly acknowledge or clear any third-party content in this assessment material, OCR will be happy to correct its mistake at the earliest possible opportunity.

For queries or further information please contact The OCR Copyright Team, The Triangle Building, Shaftesbury Road, Cambridge CB2 8EA.

OCR is part of the Cambridge Assessment Group; Cambridge Assessment is the brand name of University of Cambridge Local Examinations Syndicate (UCLES), which is itself a department of the University of Cambridge.